

Stitt's Calculations

AP Homework  
9-23-09

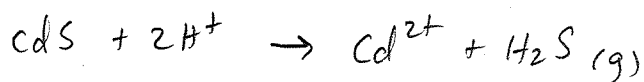
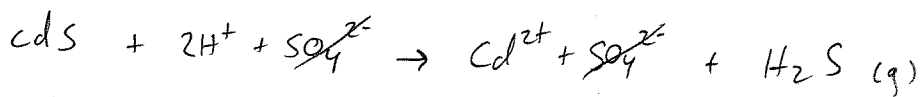
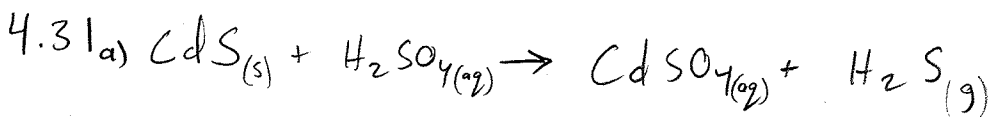
Ch 4 # 7, 20, 27, 31, 39, 42

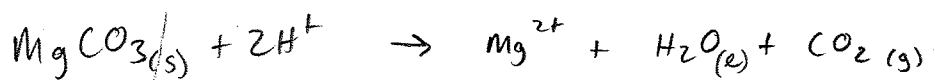
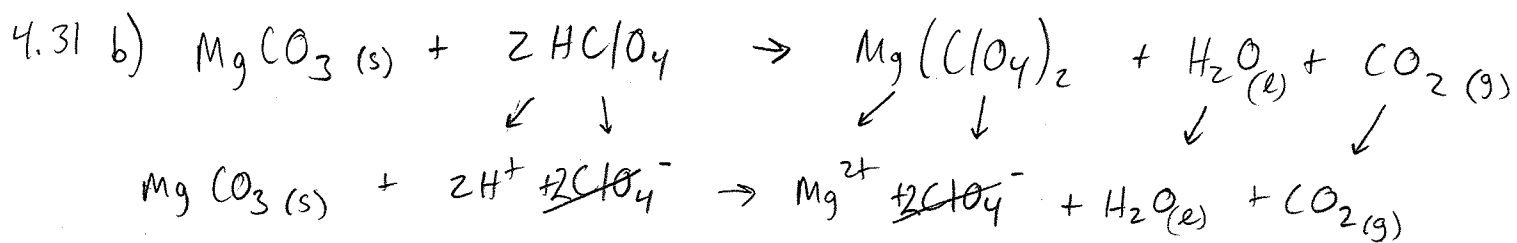
- 4.7. a) Nonelectrolyte - No ions formed  
b) weak - Few ions form  
c) strong - All become ions

4.20

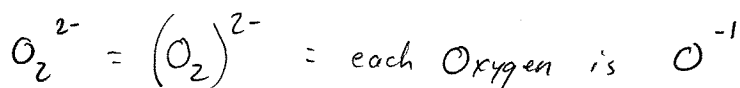
	$\text{Ba}(\text{NO}_3)_2$	$\text{NaCl}$
$\text{AgNO}_3$	No ppt	$\text{AgCl}$ ppt
$\text{CaCl}_2$	No ppt	No ppt
$\text{Al}_2(\text{SO}_4)_3$	$\text{BaSO}_4$ ppt	No ppt.

- 4.27 a) weak Acid = weak Electrolyte  
b) No ions = molecular = Non Electrolyte  
c)  $\text{NH}_3$  = weak base = weak Electrolyte  
d.  $\text{KClO}_3$  = Group I ionic compound = strong Electrolyte  
e.  $\text{Ca}(\text{NO}_3)_2$  = Very soluble ionic = strong Electrolyte.





- 4.39)
- |       |       |
|-------|-------|
| a) +6 | d) +1 |
| b) +4 | e) 0  |
| c) +7 | f) -1 |



- 42)
- Acid/Base
  - Oxidation - Reduction (REDOX) Fe = reduced, C = oxidized
  - precipitation reaction
  - Oxidation - Reduction (REDOX) Zn = oxidized, N = reduced